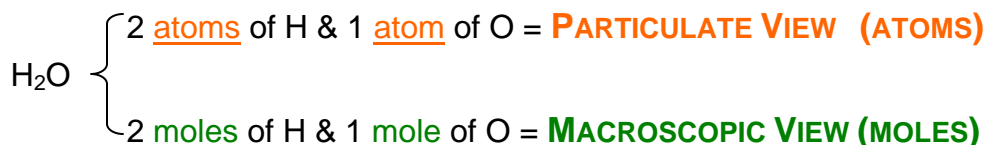


EMPIRICAL FORMULA – Lowest Numerical Ratio

Steps:

1. Find moles of each element in the compound
2. Write formula with mole subscripts
3. Divide all by smallest number
4. Round off or multiply



IMPORTANT MEMORY WORK: (for Chapter 3 and beyond)

- **STRONG ACIDS** (SE – STRONG ELECTROLYTES): [EVERYTHING ELSE IS WEAK]
 - o HClO_3 , HClO_4 , HCl
 - o HBr
 - o HI
 - o HNO_3 , H_2SO_4
- **SOLUBLE (STRONG) BASES** (STRONG HYDROXIDES): [EVERYTHING ELSE IS WEAK]
 - o H
 - o Li Be
 - o Na Mg
 - o K Ca
 - o Rb Sr
 - o Cs Ba
 - o Fr Ra

} from the periodic table
- **7 ELEMENTS EXIST AS DIATOMIC ELEMENTS:**
 1. H_2
 2. N_2
 3. O_2
 4. F_2
 5. Cl_2
 6. Br_2
 7. I_2